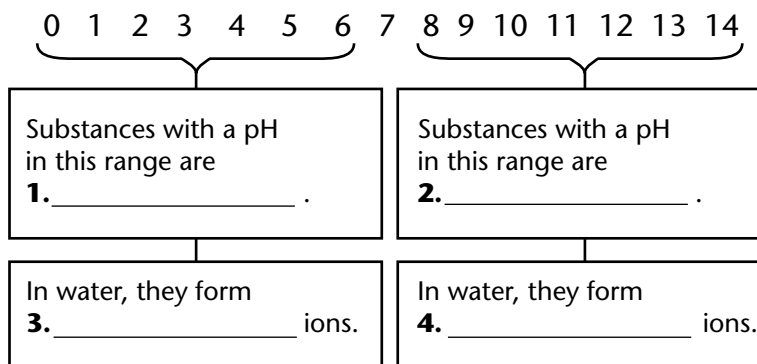


**Acids, Bases, and Solutions** ▪ *Review and Reinforce*

# Acids and Bases in Solution

## Understanding Main Ideas

Complete the concept map shown below and answer the following questions on a separate sheet of paper.



5. What is the difference between a strong acid and a weak acid?
6. What is the difference between a strong base and a weak base?
7. Which solution has a greater concentration of hydrogen ions ( $H^+$ ), a solution with a pH of 3 or one with a pH of 7? Explain.
8. What are the products formed when a base reacts with an acid?
9. What is the pH of a neutral solution?

## Building Vocabulary

Match each term with its definition by writing the letter of the correct symbol or definition on the line beside the term in the left column.

- |                        |  |
|------------------------|--|
| ___ 10. hydroxide ion  | a. ionic compound that can form from the reaction of an acid with a base           |
| ___ 11. pH scale       | b. reaction between an acid and a base   |
| ___ 12. hydrogen ion   | c. series of numbers that indicates the concentration of hydrogen ions in solution |
| ___ 13. neutralization | d. $H^+$   |
| ___ 14. salt           | e. $OH^-$  |